Honors Chemistry – Unit 3 Review

1. The electrons in the highest occupied energy level of an atom are called the 2. The	alactrons
3. Oxygen atoms attain a stable electron configuration by	ration of a noble gas
4. Ionic compounds are composed of	ation of a noble gas.
repeating 3D crystal structure. This structure makes these compounds.	are arranged in a
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5. Determine the number of valence electrons in each of the following and then draw a Lewis do str a. Magnesium b. Chlorine c. Hydrogen d. Neon 6. Write the electron configuration for the following: a. Calcium ion b. Strontium ion c. Sulfur ion d. Fluorine ion 7. Which of these is not an ionic compound? a. KF b. Na ₂ SO ₄ d. Na ₂ O 8. Using Lewis Dot diagrams, show how an ionic bond of sodium oxide is formed. 9. Describe how a metallic bond is formed. Chapter 8 – Covalent Bonding 1. Covalent bonds occur between of electrons. 2. How many electrons are shared in the following bonds? a. Single covalent bond b. Double covalent bond c. Triple covalent bond 3. For the following compounds – CF ₄ , CO ₂ , NH ₃ , N ₂ , CO, SF ₆ , BF ₃ , CH ₂ Cl ₂ , H ₂ O – do the followin a. Draw the Lewis Dot structure b. Determine the number of lone pair electrons on the central atom c. Determine the number of atoms bonded to the central atom d. Indicate the VSEPR geometry for each molecule e. Determine if the bonds are nonpolar covalent or polar covalent f. Determine if the molecule is nonpolar or polar g. Determine the type(s) of intermolecular attractions (dispersion, dipole, hydrogen bonding h. Calculate the bond dissociation energy for all bonds in the molecules 4. Order the types of bonds from strongest to weakest: London dispersion forces, hydrogen bonds, of ionic bonds, covalent bonds. 5. Which of these molecules can form a hydrogen bond with a water molecule? a. N ₂ b. NH ₃ d. CH ₄ 6. Draw all forms of the Lewis structure of the nitrate polyatomic ion. Chapter 9 – Chemical Names & Formulas	, WHOH
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1. nitrogen trifluoride 9. phosphorus triiodide 17. H ₂ SC)3
2. barium phosphide 10. disulfur decafluoride 18. CCl ₄	
• •	um iodide
	obromic acid
5. Cu(OH) ₃ 13. aluminum phosphate 21. brom	
6. ammonium carbonate 14. magnesium perchlorate 22. SrCl ₂	
7. carbonic acid 15. iron (III) sulfide 23. PbS	•
8. HCl 16. dinitrogen monoxide 24. dinitr	